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Answer Key Chapter 9 - Chemistry 2e | OpenStax. 1. The cutting edge of a knife that has been sharpened has a smaller surface area than a dull knife. Since pressure is force per unit area, a sharp knife will exert a higher pressure with the same amount of force and cut through material more effectively. 3.

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zombiekiller_sarah PLUS. CHAPTER 9-Mastering Chemistry. d. strong electrolyte. c. weak electrolyte. e. the amount of ion that carries 1 mol... e. saturated. In water, a substance that ionizes completely in solution is c... In water, a substance that partially ionizes in solution is ca... An equivalent is...

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9.1. Write the formulas for the following coordination compounds:
(i) Tetraammineaquacobalt (III) chloride. (ii) Potassium tetracyanonickelate (II) (iii) Tris (ethane-1,2-diamine) chromium (III) chloride. (iv) Amminebromidochloridonitrito-N- platinum (II) (v) Dichloridobis (ethane-1,2-diamine) platinum (IV) nitrate.

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Chemistry (4th Edition) Burdge, Julia Publisher McGraw-Hill Publishing Company ISBN 978-0-07802-152-7

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Answer: (a) Primary valency is ionisable. Question 6. Ammonia acts as a very good ligand but ammonium ion does not form complexes because.

(a) NH_3 is a gas while NH_4^+ is in liquid form. (b) NH_3 undergoes sp^3 hybridisation while NH_4^+ undergoes $\text{sp}^3 \text{d}$ hybridisation. (c) NH_4^+ ion does not have any lone pair of electrons.

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about the IUPAC naming of the chemical compounds.

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Here we have given Plus One Chemistry Chapter Wise Questions and Answers Chapter Chapter 9 Hydrogen. Kerala Plus One Chemistry Chapter Wise Questions and Answers Chapter 9 Hydrogen. Very Short Answer Type Questions (Score 1) Question 1. Producer gas is a mixture of

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As we noted in Chapter 9.2 the heat released by a reaction carried out at constant volume is identical to the change in internal energy (ΔE) rather than the enthalpy change (ΔH); ΔE is related to ΔH by an expression that depends on the change in the number of moles of gas during the reaction. The difference between the heat flow measured at constant volume and the enthalpy change is usually quite small, however (on the order of a few percent).

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