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Introduction, Multiple Choice Practice

Problems Chemistry Chemistry 12.1 What are

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9. Rank the acids by increasing acid strength: 10. Explain in terms of the partial

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charge on hydrogen why NaOH is a base, HClO is a weak acid and HClO₄ is a strong acid.

11. Why is HCl a strong acid and HClO a weak acid? 12. Why are HCl and HClO₄ both strong acids? 13.

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WS-1. Guide to Acids and Bases. Strong acids ionize 100% by the following reactions: a) $\text{HClO}_4 (\text{aq}) + \text{H}_2\text{O} (\text{l}) \rightarrow \text{H}_3\text{O}^+ (\text{aq}) + \text{ClO}_4^- (\text{aq})$ b) $\text{H}_2\text{SO}_4 (\text{aq}) + \text{H}_2\text{O} (\text{l}) \rightarrow \text{H}_3\text{O}^+ (\text{aq}) + \text{HSO}_4^- (\text{aq})$

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following acids or bases in water. Make sure you indicate whether this should be considered an equilibrium or not. a) Sulfuric acid (H_2SO_4) b) Perchloric acid (HClO_4) c)

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Potassium hydroxide d) Ammonia e) Nitrous acid (HNO_2) 3. Find the conjugate acid/base pairs in the following: a. $\text{HNO}_2 + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{NO}_2^-$ b. $\text{NH}_3 + \text{H}_2\text{O} \rightleftharpoons \text{NH}_4^+$

~~Acids and Bases Worksheet 1 — Supplemental Instruction UCF~~

10. Explain in terms of the partial charge on hydrogen why NaOH is a base, HClO is a weak acid and HClO_4 is a strong acid. 11. Why is HCl a strong acid and HClO a weak acid? 12. Why are HCl and HClO_4 both strong acids? 13. For each of the reactions below, classify the reactants as an acid or a base and the

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products as the conjugate acid or ...

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~~Acids and Bases Worksheet 1~~

Compounds formed from combinations of acids
and bases taste salty and are referred to in
chemistry as salts. Sweet compounds have
characteristics of both acids and bases in

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the same molecule. We will explore the relationship between molecular structure and acids-bases, and consider water solutions of acids and bases.

~~Introduction to Acids and Bases (Worksheet)~~
~~Chemistry ...~~

1. Acids turn blue litmus red while bases turn red litmus blue. 2. According to the Arrhenius definition, an acid is a substance that produces H^+ (hydrogen ions) in dilute aqueous solution. For example: $HCl(aq) \rightarrow H^+(aq) + Cl^-(aq)$ 3. According to the Arrhenius definition, a base is a substance that

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produces OH⁻ (hydroxide ions) in dilute aqueous solution.

~~U8LM1B WS Acids and Bases Basics Name: KEY~~
Acids, Bases, and Solutions ANSWER KEY Acids, Bases, and Solutions Describing Acids and Bases Review and Reinforce 1 Acids and bases worksheet 1 answer key. Sour 2. Bitter 3. Corrosive to magnesium, zinc, and iron; eats them away and produces bubbles of hydrogen gas 4. Doesn't react with metals 5. Produces carbon dioxide 6 Acids and bases worksheet 1 answer key.

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~~Acids And Bases Worksheet 1 Answer Key~~

NaOH is a Brønsted-Lowry base because it accepts H^+ ions from acids. NaOH is a Lewis base because the lone pairs on the hydroxide ion can be donated to other compounds. 3)

When hydrogen chloride reacts with ammonia, ammonium chloride is formed. Write the equation for this process, and indicate which of the reagents is the Lewis acid and which is the Lewis base. $HCl + NH_3 \rightarrow NH_4Cl$. NH_3 is a Lewis base because it uses its lone pair electrons to pull a hydrogen atom from hydrochloric acid.

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of acids and bases, write equations for the following acid-base reactions and indicate each conjugate acid-base pair: a) $\text{HNO}_3 + \text{OH}^-$ ($\text{H}_2\text{O} + \text{NO}_3^-$ - HNO_3 and NO_3^- make one pair. OH^- and H_2O make the other. b) $\text{CH}_3\text{NH}_2 + \text{H}_2\text{O}$ ($\text{CH}_3\text{NH}_3^+ + \text{OH}^-$ - CH_3NH_2 and CH_3NH_3^+ make one pair

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Acids and bases Acids, bases and alkalis are found in the laboratory and at home. Acids and bases can neutralise each other. A base that can dissolve in water is also called an alkali.

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